



Name: M. Umer Farooq, Quiz Subject:

Chemistry

Time Remaining: 45/45 (Minutes)

Q.1

Test 5 Solids

CHEMISTRY NMDCAT

A crystal 'x' has very high melting point and is totally insoluble in water and does not conduct electricity is likely to be _____ crystal:

- a. Ionic
- b. Covalent
- c. Molecular
- d. Metallic

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Correct Answer:

- A
- B
- C
- D

Next



Time Remaining: 44/45 (Minutes)

Q.2

Test 5 Solids

CHEMISTRY NMDCAT

Dry Ice Is solid:

- a. SO₂
- b. CO₂
- c. CO
- d. O₂

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Correct Answer:

- A
- B
- C
- D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.3

Test 5 Solids

CHEMISTRY NMDCAT

The inter ionic distance is maximum in the crystal lattice of

- a. LiCl
- c. RbCl

- b. NaCl
- d. KCl

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Correct Answer:

- A
- B
- C
- D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.4

Test 5 Solids

CHEMISTRY NMDCAT

Among the properties which property correctly match with metallic solids:

- i. They show metallic lusters
 - ii. Their conductance decreases with the rise of temperature
 - iii. Malleable and ductile
 - iv. Cubic and hexagonal close packing
- a. i,ii and iv
 - b. i,ii and iii
 - c. i ,ii, iii and iv
 - d. I,iii and iv

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Correct Answer:

- A
- B
- C
- D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.5

Test 5 Solids

CHEMISTRY NMDCAT

Which statement about covalent solids is incorrect?

- a. They contain a network of atoms
- b. They have high melting points
- c. They are very hard and greater energy is required to break them
- d. Their volatility is very high

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Correct Answer:

A B C D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.6

Test 5 Solids

CHEMISTRY NMDCAT

Covalent crystals are bad conductor of electricity due to absence of free electrons and ions except:

- a. Silicon Carbide
- b. Graphite
- c. Cadmium iodide
- d. Born nitride

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Correct Answer:

- A
- B
- C
- D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.7

Test 5 Solids

CHEMISTRY NMDCAT

The covalent crystals having giant molecules like diamond and silicon carbide are:

- a. Soluble in all the solvents
- b. Insoluble in all the polar solvents only
- c. Soluble in all the non-polar solvents only
- d. Insoluble in all the solvents

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Correct Answer:

A B C D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.8

Test 5 Solids

CHEMISTRY NMDCAT

In diamond each carbon atom is:

- a. sp-hybridized
- b. sp^2 -hybridized
- c. sp^3 -hybridized
- d. unhybridized

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Correct Answer:

- A
- B
- C
- D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.9

Test 5 Solids

CHEMISTRY NMDCAT

The overall structure of diamond looks like:

- a. Body centered cubic
- b. Face centered cubic
- c. End centered
- d. none of given

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Correct Answer:

A B C D

Next

Back



Time Remaining: 44/45 (Minutes)

Q.10

Test 5 Solids

CHEMISTRY NMDCAT

Which one of the following pairs contains polar molecular solids?

- a. Iodine and Sugar
- b. Carbon dioxide and Ice
- c. Phosphorus and Carbon dioxide
- d. Sugar and Ice

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Correct Answer:

- A
- B
- C
- D

Next

Back



Time Remaining: 44/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Valence bond theory treat metallic bond as:

- a. A coordinate covalent bond
- b. A localized covalent bond
- c. A de-localized covalent bond
- d. Van der Waal's force

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Page 18 of 25 Page Last Update: 10:58 AM

Correct Answer:



Next

Back



Time Remaining: 44/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Silicon carbide forms

- a. ionic solid crystal
- b. covalent crystal
- c. molecular crystal
- d. b, c

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Page 18 of 25 | Page Last Updated: 10/10/2023

Correct Answer:



Next

Back



Time Remaining: 43/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Substance having m.p higher than 500°C and insoluble in H₂O and organic solvent and conductor in solid and phase:

- a. Copper
- b. Sodium chloride
- c. Silica
- d. Cell

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Page 18 of 20 Page Last Update

Correct Answer



Next

Back



Time Remaining: 43/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The crystalline solid which show very slow chemical reactions

- a. ionic
- c. molecular
- c. covalent
- d. b, c

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Next

Back



Time Remaining: 45/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The arrangement ABC, ABC is referred as

- a. cubic close packing
- b. octahedral closed packing
- c. hexagonal closed packing
- d. tetrahedral closed packing

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Page 1 of 1 Page Last Updated: 10/10/2023

Correct Answer:



Next

Back



Time Remaining: 45/45 (Minutes)

Q10

Test 5 Solids

CHEMISTRY NMDCAT

Which statement is not true about metallic solid?

- a. They are ductile and malleable
- b. They are conductor of electricity
- c. their conductivity increases by increasing temperature
- d. they are lustrous

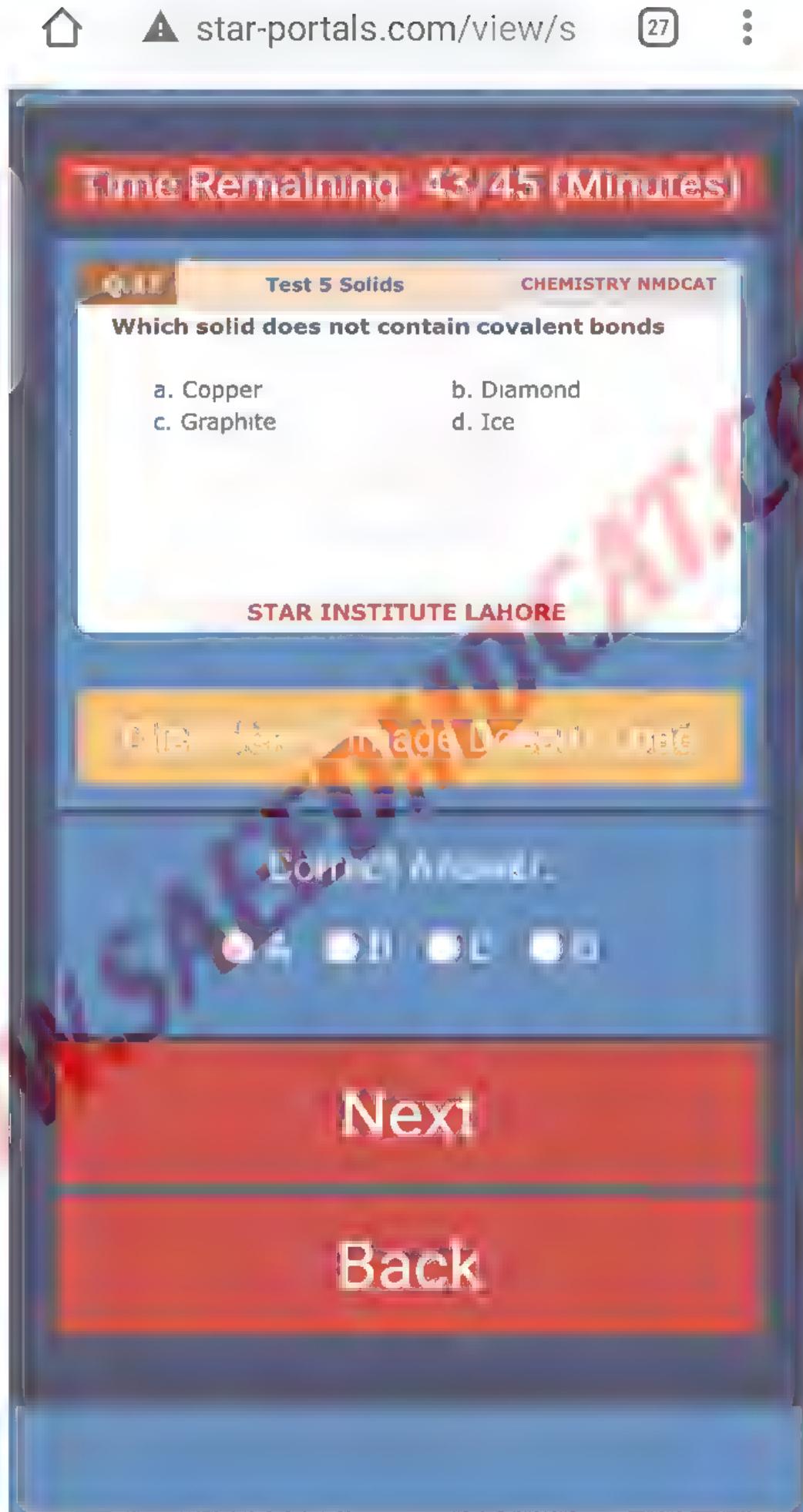
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P B I S M Page 1 of 1 Date _____

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Next

Back





Time Remaining: 45/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The branch of science which deals with the study of the structure of crystal is called

- a. Chymography
- b. Crystallography
- c. Chromatography
- d. Industrial chemistry

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Page 1 of 1 Page Last Updated: 10:53 AM

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Next

Back



Time Remaining: 46/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Glass may begin to crystallize by a process called

- a. annealing
- b. etching
- c. distillation
- d. none of these

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Page 18 of 20 | Page Last Updated: 10/10/2023

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Next

Back



Time Remaining: 45/45 (Minutes)

Q30

Test 5 Solids

CHEMISTRY NMDCAT

The crystalline part of other wise amorphous solids is called

- a. crystal system
- b. crystallite
- c. crystal lattice
- d. allotrope

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Page 1 of 1 | Page 1 of 1

Correct Answer:

Next

Back



Time Remaining: 45/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Which substance has diffused melting point

- a. crystalline solid
- b. amorphous solid
- c. metallic solid
- d. covalent solid

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Page 1 of 1 | Page Last Updated: 10/10/2023

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Next

Back



Time Remaining: 43:45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Which has the strongest bonding in the solid state?

- a. Hydrogen Chloride (HCl)
- b. Chlorine (Cl₂)
- c. Xenon(Xe)
- d. Sodium Chloride (NaCl)

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Page 1 of 1 | Page Last Updated: 10:54 AM

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Next

Back



Time Remaining: 45 Minutes



Test 5 Solids

CHEMISTRY NMDCAT

When substance moves from a solid to a liquid state all of the following changes occur except

- a. Molecules become more disordered
- b. K.E of the molecules decreases
- c. Intermolecular forces become weaker
- d. Molecule become further separated

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Page 1 of 1 Page Last Updated: 10/10/2016

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Next

Back



Time Remaining: 45 Minutes



Test 5 Solids

CHEMISTRY NMDCAT

When the atoms of the third layer are arranged in such a way that they directly lie above the atoms of the first layer then this arrangement is called

- a. ABAB (hexagonal)
- b. ABCABC (Cubic)
- c. Orthorhombic
- d. Rhombohedral

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Page 18 of 20

Correct Answer:



Next

Back



Time Remaining: 45 Minutes



Test 5 Solids

CHEMISTRY NMDCAT

Which one is a conductor but is not malleable?

- a. Iron
- b. Graphite
- c. Silver
- d. Platinum

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Page 1 of 10 | Page Last Updated: 10/10/2023

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Next

Back



Time Remaining: 42:45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

A malleable solid is one which can be

- a. Converted into wires
- b. Converted into thin sheets
- c. Melted easily
- d. All of the above

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Page 18 of 20 Page Last Updated: 10/10/2016

Correct Answer:



Next

Back





Time Remaining: 45 Minutes



Test 5 Solids

CHEMISTRY NMDCAT

Buckyballs is an allotropic form of

- a. Sulphur
- b. Carbon
- c. Silica
- d. Tin

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Page 1 of 1 Page Last Updated: 10/10/2016

Correct Answer:



Next

Back

Chemistry

Time Remaining 42/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The combined state of the metal is called?

- a. Solid
- b. Metal
- c. Both A & B
- d. Minerals

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Next

Back



Time Remaining: 42/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

A solid, liquid, and gas can exist together at the

- a. sublimation point
- b. triple point
- c. boiling point
- d. freezing point

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Page 18 of 25 | Page Last Updated: 10/10/2016

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Next

Back



Time Remaining: 45 Minutes



Test 5 Solids

CHEMISTRY NMDCAT

The glue comes under the example of

- a. crystalline solids
- b. amorphous solids
- c. simple solids
- d. compound solids

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Page 1 of 10 Page Last Updated: 10/10/2023

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Next

Back



Time Remaining: 42/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The ions of the ionic crystals become free when it is in

- a. solid state
- b. compound state
- c. molten state
- d. none of above

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Page 1 of 1

Correct Answer:



Next

Back



Time Remaining: 45 Minutes



Test 5 Solids

CHEMISTRY NMDCAT

The crystal which is used to study Avogadro's number is called

- a. LiF
- b. NaI
- c. NaCl
- d. KCl

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Page 1 of 10

Correct Answer:



Next

Back



Time Remaining: 42:45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The bond distance in the iodine molecule in solid iodine is

- a. 271.5pm
- b. 266.6pm
- c. 250pm
- d. 230pm

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Page 1 of 1 | Page Last Updated: 10:45 AM

Correct Answer:



Next

Back



Time Remaining: 41/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The energy which is released when 1 mole of the ionic crystal is formed is known as

- a. lattice energy
- b. heat energy
- c. molar energy
- d. kinetic energy

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Page 18 of 25 | Page Last Updated: 10/10/2023

Correct Answer:



Next

Back



Time Remaining: 41/45 (Minutes)

Q30

Test 5 Solids

CHEMISTRY NMDCAT

The isomorph of sodium fluoride is

- a. chlorine
- b. magnesium oxide
- c. sodium
- d. Sulphur

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Page 18 of 25 | Page Last Updated: 10/10/2023

Correct Answer:

Next

Back



Time Remaining: 41/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Lattice points have another name which is called lattice

- a. sites
- b. arrangements
- c. circles
- d. array

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Page 18 of 25 Page Last Update

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Next

Back



Time Remaining: 41/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The electrical properties of solid iodine include that it is

- a. non conductor
- b. conductor
- c. poor conductor
- d. moderate conductor

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Page 18 of 20

Correct Answer:



Next

Back





Time Remaining: 41/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Electrical and thermal properties for some crystalline solids depend upon

- a. surface
- b. area
- c. direction
- d. density

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Page 18 of 20 Page Last

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Next

Back



Time Remaining: 41/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

Covalent crystalline solids are soluble in

- a. polar solvents
- b. non polar solvents
- c. water
- d. Normal saline

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Page 18 of 25 | Page Last Updated: 10/10/2023

Correct Answer:



Next

Back



Time Remaining: 41/45 (Minutes)



Test 5 Solids

CHEMISTRY NMDCAT

The depressions among two layers of metals are also known as

- a. voids
- b. holes
- c. window
- d. shuttle

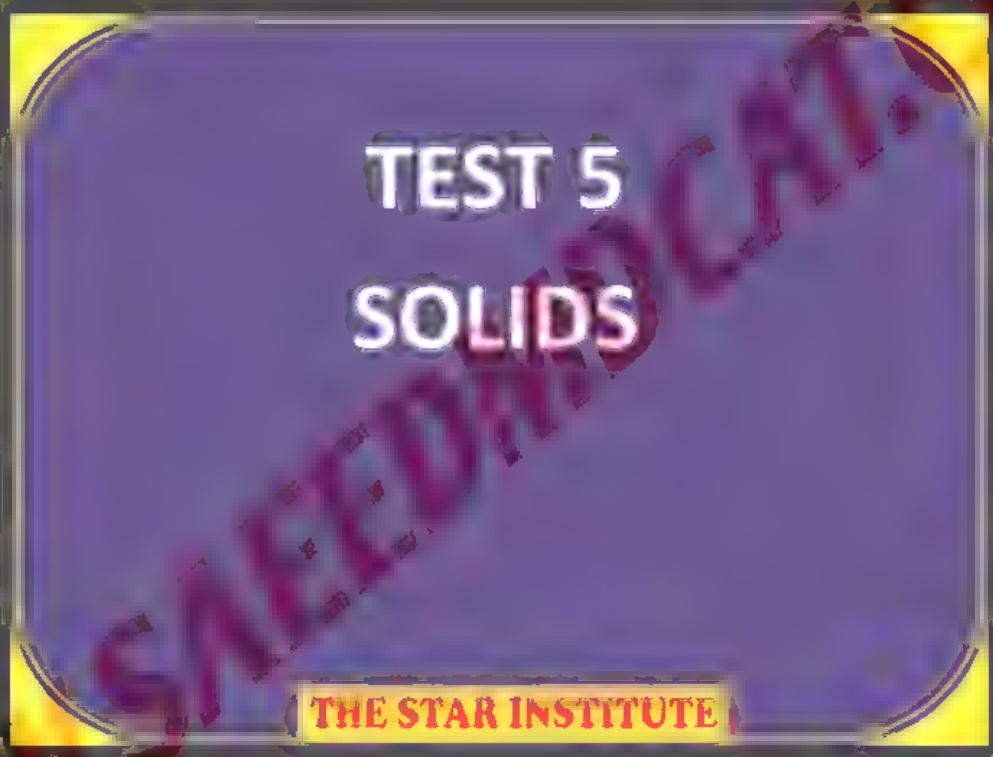
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Page 18 of 25 | Page Last Updated: 10/10/2023

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Submit Quiz

Back



Q. 1

A crystal 'x' has very high melting point and is totally insoluble in water and does not conduct electricity is likely to be _____ crystal:

- a. Ionic
- b. Covalent
- c. Molecular
- d. Metallic

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Q. 2

Dry ice is solid:

- a. SO_2
- b. CO
- c. CO_2
- d. O_2

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Q. 3

The inter ionic distance is maximum in the crystal lattice of

- a. LiCl
- b. NaCl
- c. RbCl
- d. KCl

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Q. 4

Among the properties which property correctly match with metallic solids:

- i. They show metallic lusters
 - ii. Their conductance decreases with the rise of temperature
 - iii. Malleable and ductile
 - iv. Cubic and hexagonal close packing
- a. i,ii and iv
 - b. i,ii and iii
 - c. i,ii, iii and iv
 - d. i,iii and iv

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Q. 5

Which statement about covalent solids is incorrect?

- a. They contain a network of atoms
- b. They have high melting points
- c. They are very hard and greater energy is required to break them
- d. They volatility is very high

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Q. 6

Covalent crystals are bad conductor of electricity due to absence of free electrons and ions except:

- a. Silicon Carbide
- b. Graphite
- c. Cadmium iodide
- d. Boron nitride

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Q. 7

The covalent crystals having giant molecules like diamond and silicon carbide are:

- a. Soluble in all the solvents
- b. Insoluble in all the polar solvents only
- c. Soluble in all the non-polar solvents only
- d. **Insoluble in all the solvents**

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Q. 8

In diamond each carbon atom is:

- a. sp- hybridized
- b. sp^2 -hybridized
- sp^3 hybridized**
- d. unhybridized

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Q. 9

The overall structure of diamond looks like:

- a. Body centered cubic
- b. Face centered cubic
- c. End centered
- d. none of given

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Q. 10

Which one of the following pairs contains polar molecular solids?

- a. Iodine and Sugar
- b. Carbon dioxide and Ice
- c. Phosphorus and Carbon dioxide
- d. Sugar and Ice

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Q. 11

Valence bond theory treat metallic bond as:

- a. A coordinate covalent bond
- b. A localized covalent bond
- c. A de-localized covalent bond
- d. Van der Waal's force

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Q. 12

Silicon carbide forms

- (a) ionic solid crystal
- (b) covalent crystal
- (c) molecular crystal
- (d) b, c

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Q. 13

Substance having m.p higher than 500°C and insoluble in H₂O and organic solvent and conductor in solid and phase:

- a. Copper
- b. Sodium chloride
- c. Silica
- d. Cell

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Q. 14

The crystalline solid which show very slow chemical reactions

- (a) ionic
- (b) covalent
- (c) molecular
- (d) b, c

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Q. 15

The arrangement ABC, ABC is referred as

- (a) simple cubic packing
- (b) octahedral closed packing
- (c) hexagonal closed packing
- (d) tetrahedral closed packing

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Q. 16

Which statement is not true about metallic solid?

- (a) They are ductile and malleable
- (b) They are conductor of electricity
- (c) their conductivity increase, as temperature increases
- (d) they are lustrous

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Q. 17

Which solid does not contain covalent bonds

- (a) Copper
- (b) Diamond
- (c) Graphite
- (d) Ice

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Q. 18

The branch of science which deals with the study of the structure of crystal is called

- (a) Chymography
- (b) Crystallography**
- (c) Chromatography
- (d) Industrial chemistry

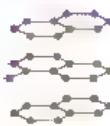
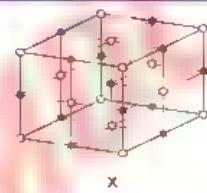
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Q. 19

The diagram shows part of the lattice structures of solids X and Y.

What are the types of bonding present in X and Y?



	X	Y
A	covalent	metallic
B	ionic	covalent
C	ionic	metallic
D	metallic	ionic

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Q. 20

The crystalline part of otherwise amorphous solids is called

- (a) crystal system
- (b) crystal
- (c) crystal lattice
- (d) allotrope

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Q. 21**Which substance has diffused melting point**

- (a) crystalline solid (b) amorphous solid
- (c) metallic solid (d) covalent solid

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Q. 22

Which has the strongest bonding in the solid state?

- A. Hydrogen Chloride (HCl)
- B. Chlorine (Cl₂)
- C. Xenon(Xe)
- D. Sodium Chloride (NaCl)

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Q. 23

When substance moves from a solid to a liquid state all of the following changes occur except

- A. Molecules become more disordered
- B. Intermolecular distance increases
- C. Intermolecular forces become weaker
- D. Molecule become further separated

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Q. 24

When the atoms of the third layer are arranged in such a way that they directly lie above the atoms of the first layer then this arrangement is called

- A. ABAB Hexagonal
- B. ABCABC (Cubic)
- C. Orthorhombic
- D. Rhombohedral

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Q.25

Which one is a conductor but is not malleable?

- A. Iron
- B. Gold
- C. Silver
- D. Platinum

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Q. 26

A malleable solid is one which can be

- A. Converted into wire
- B. Expanded fit in holes
- C. Melted easily
- D. All of the above

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Q. 27

Buckyballs is an allotropic form of

- A. Sulphur
- B. Carbon
- C. Silica
- D. Tin

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Q. 28

The combined state of the metal is called?

- A. Solid
- B. Metal
- C. Both A & B
- D. Metalloid

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Q. 29

- A solid, liquid, and gas can exist together at the
- A. sublimation point
 - B. triple point
 - C. boiling point
 - D. freezing point

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Q. 30

The glue comes under the example of

- A. crystalline solids
- B. amorphous solids
- C. simple solids
- D. compound solids

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Q. 31

The ions of the ionic crystals become free when it is in

- a) solid state
- b) compound state
- c) liquid state
- d) none of above

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Q. 32

The crystal which is used to study Avogadro's number is called

- a) Li
- b) NaI
- c) NaCl
- d) KCl

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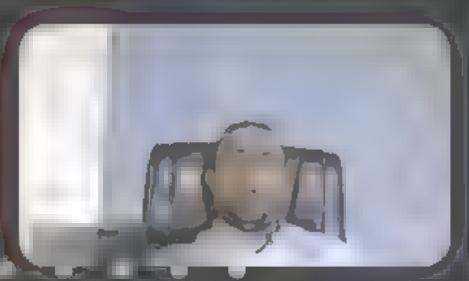


Q. 33

The bond distance in the iodine molecule in solid iodine is

- a) 277 pm
- b) 266.6pm
- c) 250pm
- d) 230pm

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Q. 34

The energy which is released when 1 mole of the ionic crystal is formed is known as

- a) lattice energy
- b) heat energy
- c) molar energy
- d) kinetic energy

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Q. 35

The isomorph of sodium fluoride is

- a) chlorine
- b) magnesium
- c) sodium
- d) Sulphur

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Q. 36

Lattice points have another name which is called lattice

- a) dots
- b) arrangements
- c) circles
- d) array

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Q. 37

The electrical properties of solid iodine include that it is

- a) non conductor
- b) conductor
- c) poor conductor
- d) moderate conductor

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Q. 38

Electrical and thermal properties for some crystalline solids depend upon

- a) surface
- b) area
- c) direction
- d) density

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Q. 39

Covalent crystalline solids are soluble in

- a) polar solvents
- b) non polar solvents
- c) water
- d) Normal saline

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Q. 40

The depressions among two layers of metals are also known as

- a) voids
- b) holes
- c) window
- d) shuttle

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